Electrochemical Cell Construction (ie. Making Batteries)

Use the following link to make your Cell,

<http://www.kentchemistry.com/moviesfiles/Units/Redox/voltaiccelll20.htm>

1. Construct a cell with silver and copper electrodes. Use silver nitrate solution in the beaker with silver and copper nitrate in the beaker with copper.
2. Turn on the voltmeter and a positive voltage reading should show if you set-up cell correctly.
3. Which electrode (silver metal or copper metal) is being oxidized? How do you know?
4. Which electrode is being reduced? How do you know?
5. Which direction do the electrons flow in the wire? Suggest a reason as to why this happens?
6. What is the voltage produced? Could you run your Television remote with this amount of voltage?
7. Try to find another combination that produces more voltage than silver and copper. Is there one? If so, which combination and what is the total voltage?