

1. On the left side of the compound, indicate if the following compound is ionic(I) or covalent(C) and then write the formulas for the following compounds:

- a. calcium iodide Ca I₂
- b. silver nitride Ag₃N
- c. lead (IV) fluoride Pb F₄
- d. iron (III) chloride FeCl₃
- e. diselenium tetracarbide Se₂C₄
- f. silicon dibromide Si Br₂
- g. barium nitrate Ba (NO₃)₂
- h. zinc phosphate Zn₃(PO₄)₂
- i. ammonium sulphate (NH₄)₂ SO₄
- j. nickel (II) hydroxide Ni (OH)₂

2. On the left side of the compound, indicate if the following compound is ionic(I) or covalent(C) and then write the names for the following compounds (Use roman numerals, brackets, and prefixes where appropriate):

- a. HBr hydrogen bromide
- b. AlCl₃ Aluminum chloride
- c. FeO iron (II) oxide
- d. SnO₂ tin (IV) oxide
- e. P₂O₃ diphosphorus trioxide
- f. SiF₄ silicon tetrafluoride
- g. Fe₂(CO₃)₃ iron (III) carbonate
- h. NH₄Cl ammonium chloride
- i. Ba(ClO₃)₂ barium chlorate
- j. Mn(SO₄)₂ manganese (IV) sulphate