mole practice test

Multiple Choice

1b 2b 3d 4d 5b 6b 7b 8c 9d 10a 11a

12b 13d 14d 15b 16c 17d 18c 19d 20a 21a 22d

23c 24c 25d 26d 27a 28a 29d 30d 31d 32a 33d

Written

1. Chemist has 1.71 x 10-3 moles, therefore chemist does not have enough because s/he needs 2.00 x 10-3.

2. 201845 g/year = 202 kg/year

3. 613 g CuSO4, 3.84 mol CuSO4

4. C7 H6 O2

No, I do not agree because it is a different formula than C12 H14 O2

5. Emperical formula = C2 H N O2

Ratio = 3

actual formula (molecular formula) = C6 H3 N3 O6

6. 50. L

7. 0.70 mol

8. 57.9 degrees celsius

9. 15.6 g CO

10. 1.74 x 10 24 atoms